5-1-2008

Formation of Bicyclic Pyrroles from the Catalytic Coupling Reaction of 2,5-disubstituted Pyrroles with Terminal Alkynes, Involving the Activation of Multiple C-H bonds

Chase S. Yi
Marquette University, chae.yi@marquette.edu

Jie Zhang
Marquette University

Accepted version. *Chemical Communications*, Vol. 20 (May 2008): 2349-2351. DOI.

Formation Of Bicyclic Pyrroles From The Catalytic Coupling Reaction Of 2,5-Disubstituted Pyrroles With Terminal Alkynes, Involving The Activation Of Multiple C–H Bonds†

Chae S. Yi
Department of Chemistry, Marquette University
Milwaukee, WI

Jie Zhang
Department of Chemistry, Marquette University
Milwaukee, WI

Abstract: Substituted bicyclic pyrroles are produced directly from the coupling reaction of 2,5-disubstituted pyrroles with terminal alkynes, involving the activation of multiple C–H bonds and regioselective cyclisation.

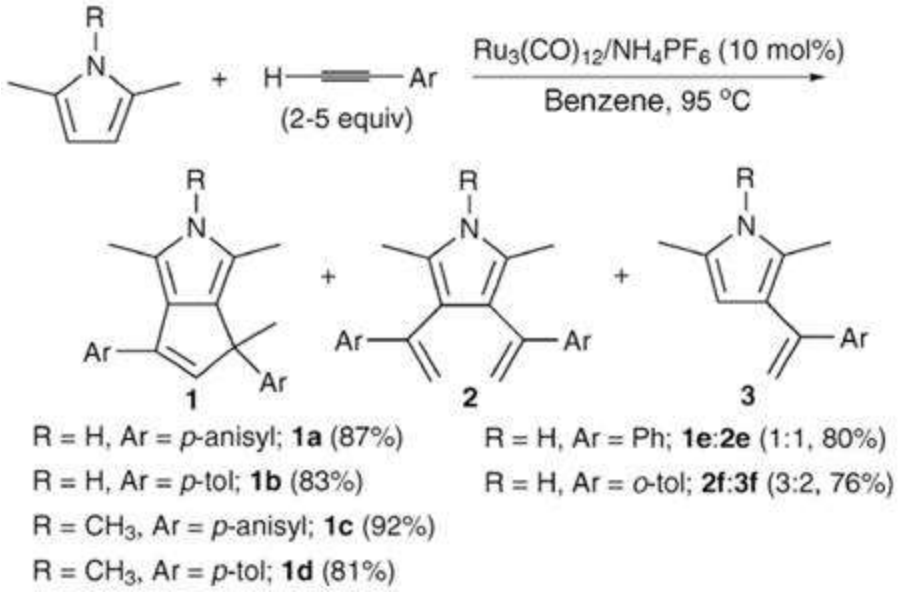
Transition metal-catalysed C–H bond activation and functionalisation reactions of nitrogen heterocyclic compounds have attracted considerable attention, in part due to their prominent role in the synthesis of natural products and pharmaceutical agents.[1] Highly regioselective catalytic C–H bond insertion reactions of nitrogen-containing aromatic compounds, such as pyridines, indoles and pyrroles, have been reported in recent years.[2] Direct oxidative coupling reactions of arene C–H bonds[3] and the C–H bond oxidative

Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.

annulation of indoles[4] have also been achieved using Cu and Pd catalysts. Despite such remarkable progress, however, catalytic C–H bond activation methods have rarely been employed for constructing nitrogen-containing heterocyclic compounds. We recently developed a new catalytic coupling reaction between arylamines and alkynes, which involved the regioselective activation of sp^2 C–H bonds to yield tricyclic quinoline products.[5] In an effort to extend the scope of catalytic C–H bond activation reactions, we have begun to explore the coupling reactions of pyrroles and indoles. This report delineates the coupling reaction between 2,5-disubstituted pyrroles and terminal alkynes, which involves multiple C–H bond activation and cyclisation steps.

Treatment of 2,5-dimethylpyrrole (1.0 mmol) with 4-ethynylanisole (2.0 mmol) in the presence of $Ru_3(CO)_{12}/NH_4PF_6$ (1: 3, 10 mol% Ru) in benzene (5 mL) at 95 °C for 36 h cleanly produced the cyclisation product, **1a** (eqn (1)). Since **1a** was found to be air sensitive, the analytically pure product was isolated by column chromatography under a nitrogen atmosphere (87% yield), and was fully characterised by both spectroscopic methods and elemental analysis.[‡] The initial catalyst activity survey showed that both $Ru_3(CO)_{12}$ and NH_4PF_6 were essential for catalytic activity. Other neutral and cationic ruthenium compounds, such as $RuCl_3{\cdot}3H_2O$, $(PPh_3)_3RuHCl$, $(PCy_3)_2(CO)RuHCl$ and $[(PCy_3)_2(CO)(MeCN)_2RuH]^+BF_4^-$, did not give any coupling products under similar reaction conditions. The analogous reaction of *N*-phenylpyrrole with 4-ethynylanisole produced a mixture of 1: 1 and 1: 2 coupling products, without forming any cyclisation product.

Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.



(1)

The coupling reaction was found to be strongly influenced by the steric and electronic nature of alkynes. In contrast to terminal alkynes with a *para*-electron-donating group, such as 4-ethynylanisole or 4-ethynyltoluene, which readily produced the cyclisation products **1a–1d**, the coupling reaction with phenylacetylene gave a 1: 1 mixture of the cyclisation and 1: 2 insertion products, **1e** and **2e**. The coupling reaction with sterically demanding 2-ethynyltoluene (2 equiv.) produced a 3: 2 mixture of the coupling products **2f** and **3f** under similar conditions. Neither arylalkynes with an electron-withdrawing group, such as 4-ethynyltrifluorotoluene or 4-fluorophenylacetylene, nor the aliphatic terminal alkynes, gave any coupling products under similar conditions. A prolonged reaction time at a higher temperature did not convert **2** or **3** into cyclisation product **1**. Instead, the cyclotrimerisation products from the homocoupling of the terminal alkynes were produced predominantly in these cases.

Since $Ru_3(CO)_{12}/NH_4PF_6$ was not particularly effective for the coupling reactions with electron-poor arylalkynes, we next surveyed the efficacy of gold catalysts to promote the formation of cyclisation products. When 2,5-dimethylpyrrole was treated with phenylacetylene (2 equiv.) in the presence of 5 mol% of $Au(PPh_3)Cl/AgOTf$ (1: 1) in benzene for 24 h, **1e** was formed exclusively, though the catalyst lost its activity after 60% conversion. Control experiments indicated that

Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.

both $Au(PPh_3)Cl$ and AgOTf were required for catalytic activity, and other selected gold compounds, such as $AuCl_3$ and $NaAuCl_4$, failed to catalyse the coupling reaction. When the $Au(PPh_3)Cl$/AgOTf (5 mol%) catalyst was treated with a 1: 1 mixture of **1e** and **2e**, **2e** was cleanly converted to **1e** to produce an 8: 1 mixture of **1e** and **2e** after 10 h at 95 °C. By using the combined catalytic system, $Ru_3(CO)_{12}/NH_4PF_6$ and $Au(PPh_3)Cl$/AgOTf, cyclisation product **1e** was obtained from the coupling reaction of 2,5-dimethylpyrrole with phenylacetylene (>95% conversion, 81% combined yield, **1e: 2e** = 85: 15). This result indicates that the gold catalyst was particularly effective in promoting the cyclisation step of the coupling reaction. While gold catalysts have been successfully utilised in C–H bond activation reactions,[6] the synergistic effect of Ru/Au catalysts is not entirely clear at the present time.

The formation of both 1: 1 and 1: 2 products suggested that product **1** is resulted from the cyclisation of 1: 2 coupling product **2**. To gain further mechanistic insights, the reaction mixture of **1e** and **2e** (1: 1) was periodically monitored by ^1H NMR at room temperature, after it had been heated at 95 °C in the presence of $Ru_3(CO)_{12}/NH_4PF_6$ (10 mol% Ru) in C_6D_6 (Fig. 1). Over time, the peaks due to **1e** at δ 6.19, as well as the NH peak at δ 6.24, increased at the expense of the peaks due to **2e** (δ 5.27 and 5.53 (CQCH_2)). The rate constant, $k_{obs}=2.1 \times 10^{-2}$ h^{-1}, of the appearance of **1e** was estimated from a pseudo first-order plot.

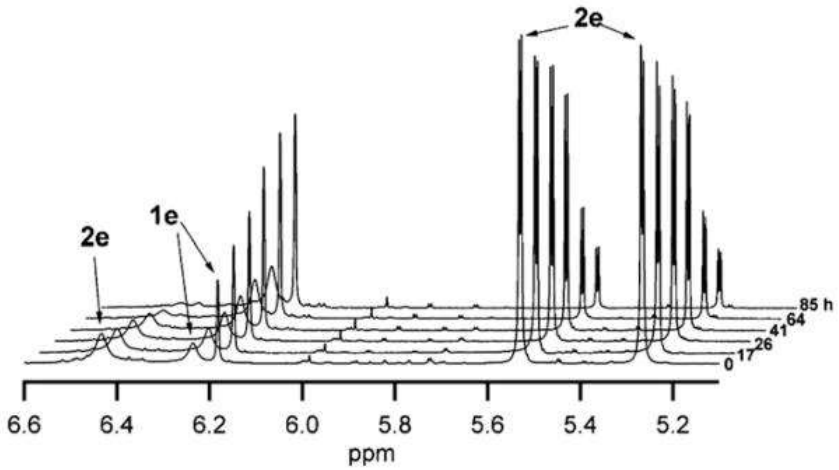
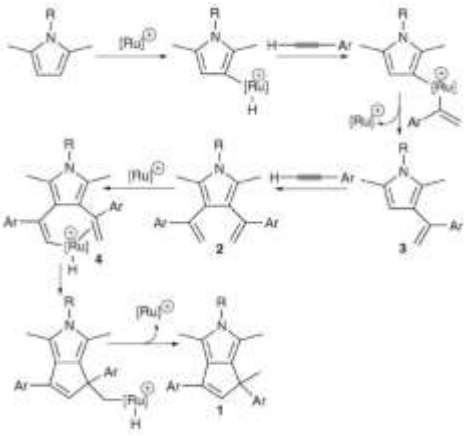


Fig. 1 Partial ^1H NMR spectra of the reaction mixture of **1e** and **2e**.

Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.

The coupling reaction of 1,2,5-trimethylpyrrole with deuterium-labelled 4-ethynylanisole-d_1 (2 equiv., >99% D) in the presence of $Ru_3(CO)_{12}/NH_4PF_6$ (10 mol% Ru) in C_6D_6 was monitored by NMR. After 1 h of heating at 95 °C, the ^{1}H NMR spectrum showed that nearly 15% of the deuterium from 4-ethynylanisole had exchanged with 35% of the β-vinyl hydrogens of the unreacted 1,2,5-trimethylpyrrole, prior to the formation of the coupling products. The product, **2a**-*d*, isolated from a preparative scale reaction of 2,5-dimethylpyrrole with 2 equivalesnts of 4-ethynylanisole-d_1, contained deuterium at both the α-methyl (33%) and vinyl (37%) positions. Also, in support of rapid H/D exchange between the two substrates, a relatively small deuterium isotope effect was observed from a separate reaction of 1,2,5-trimethylpyrrole with phenylacetylene/ phenylacetylene-d_1 when forming 1: 1 coupling product **3e**. The pseudo first-order plots for the reactions gave k_{obs} = 1.65 × 10^{-2} and 1.38 × 10^{-2} h^{-1} from phenylacetylene and phenylacetylene-d_1, respectively, which translated into k_{CH}/k_{CD} = 1.2.

These results suggest a mechanism involving sequential alkyne insertion and cyclisation steps, as outlined in Scheme 1. The sequential C–H activation and regioselective insertion of alkynes would be mediated by an electrophilic ruthenium catalyst to form 1: 2 coupling product **2**. The subsequent ruthenium-mediated vinyl C–H bond activation and cyclisation steps could be facilitated by coordination of the adjacent olefin to ruthenium *via* the formation of alkene–hydride species **4**. Cyclisation and reductive elimination would give product **1**.



Scheme 1 A possible mechanistic pathway.

Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.

In summary, the catalytic formation of bicyclic pyrroles has been achieved from the direct coupling reaction of 2,5-di-methylpyrroles with terminal alkynes. The cyclisation reaction involved three consecutive sp^2 C–H bond activation and insertion steps.

Supplementary Information

Experimental details

General information. All operations were carried out in an inert-atmosphere glove box or by using standard high vacuum and Schlenk techniques. Tetrahydrofuran, benzene, hexanes were distilled from purple solutions of sodium and benzophenone immediately prior to use. CH_2Cl_2 was distilled from CaH_2. The NMR solvents were dried from activated molecular sieves (4 Å). Pyrrole and alkyne substrates were received from commercial sources and used without further purification. The ^1H, ^2H and ^{13}C NMR spectra were recorded on a Varian Mercury 300 MHz or 400 MHz FT-NMR spectrometer. Mass spectra were recorded from a Hewlett-Packard HP5970 GC/MS spectrometer. Elemental analyses were performed at the Midwest Microlab, Indianapolis, Indiana, USA.

Typical procedure of the catalytic reaction. In a glove box, $Ru_3(CO)_{12}$ (0.03 mmol), NH_4PF_6 (0.1 mmol), 2,5-dimethylpyrrole (1.0 mmol) and an alkyne (2.0 mmol) were dissolved in 5 mL benzene solution in a medium-walled 25 mL Schlenk tube equipped with teflon stopcock and a magnetic stirring bar. The tube was sealed and was brought out of the box. The reaction tube was heated in an oil bath at 95 °C for 36-48 h. The tube was opened to air at room temperature, and the crude product mixture was analysed by GC/MS. The solvent was removed under a rotary evaporator, and the organic product was isolated by a column chromatography on silica gel (hexane/CH_2Cl_2) under nitrogen. For the combined catalyst system: 2,5-dimethylpyrrole (1.0 mmol), phenylacetylene (2.0 mmol), $Ru_3(CO)_{12}$ (0.03 mmol), NH_4PF_6 (0.1 mmol), $Au(PPh_3)Cl$ (0.05 mmol) and AgOTf (0.05 mmol) were used under otherwise same reaction conditions.

For **1a**: δ_H(300 MHz; C_6D_6) 7.60-6.83 (8 H, m, Ar), 6.25 (1 H, br s, NH), 6.13 (1H, s, C=C*H*), 3.33 (3 H, s, OCH_3), 3.32 (3 H, s, OCH_3), 2.10 (3 H, s, CH_3), 1.92 (3 H, s, CH_3), 1.87 (3 H, s, CH_3);

Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.

δ_C(75 MHz, C_6D_6) 159.6, 158.5, 141.1, 138.6, 138.1, 135.4, 130.2, 129.2, 127.8, 127.4, 117.2, 114.2, 114.1, 114.0, 54.8 (OCH_3), 54.7 (OCH_3), 50.3 (CCH_3), 24.7 (CH_3), 13.0 (CH_3), 11.8 (CH_3); *m/z* (GC-MS) 359 (M+); Found: C, 79.62; H, 7.34; N, 3.25. Calc. for $C_{24}H_{25}NO_2$: C, 80.19; H, 7.01; N, 3.90%.

For **1b**: δ_H(400 MHz; C_6D_6) 7.58-7.03 (8 H, m, Ar), 6.20 (1 H, br s, NH), 6.17 (1 H, s, C=CH), 2.15 (6 H, s, CH_3), 2.08 (3 H, s, CH_3), 1.90 (3 H, s, CH_3), 1.86 (3 H, s, CH_3); δ_C(75 MHz, C_6D_6) 142.9, 141.4, 138.9, 136.6, 135.1, 135.0, 134.7, 129.1, 127.9, 126.7, 117.0, 114.1, 50.6 (CCH_3), 24.3 (CH_3), 21.1 (CH_3), 20.9 (CH_3), 12.8 (CH_3), 11.6 (CH_3); *m/z* (GC-MS) 327 (M+); Found: C, 88.02; H, 7.62; N, 4.31. Calc. for $C_{24}H_{25}N$: C, 88.03; H, 7.70; N, 4.28%.

For **1c**: δ_H(400 MHz; C_6D_6) 7.61-6.85 (8 H, m, Ar), 6.12 (1 H, s, C=C*H*), 3.34 (6H, s, OCH_3), 2.74 (3 H, s, NCH_3), 2.12 (3 H, s, CH_3), 1.94 (3 H, s, CH_3), 1.90 (3 H, s, CH_3); δ_C(100 MHz, C_6D_6) 159.4, 158.3, 140.9, 138.7, 138.0, 134.6, 130.3, 129.2, 127.7, 126.4, 119.2, 116.2, 113.9, 113.8, 54.7 (OCH_3), 54.6 (OCH_3), 50.5 (CCH_3), 29.4 (NCH_3), 24.6 (CH3), 11.7 (CH_3), 11.0 (CH_3); *m/z* (GC-MS) 373 (M+); Found: C, 79.87; H, 7.16; N, 3.68. Calc. for $C_{25}H_{27}NO_2$: C, 80.40; H, 7.29; N, 3.75%.

For **1d**: δ_H(300 MHz; acetone-d_6) 7.42-7.04 (8 H, m, Ar), 5.93 (1 H, s, C=CH), 3.39 (3 H, s, NCH_3), 2.34 (3 H, s, CH_3), 2.26 (3 H, s, CH_3), 2.20 (3 H, s, CH_3), 2.08 (3 H, s, CH_3), 1.67 (3 H, s, CH_3); δ_C(75 MHz; acetone-d_6) 143.4, 141.0, 139.6, 137.4, 135.5, 135.2, 134.2, 129.7, 129.5, 128.3, 127.1, 126.3, 120.1, 117.0, 51.2 (CCH_3), 30.2 (NCH_3), 24.8 (CH_3), 21.2 (CH_3), 20.9 (CH_3), 11.8 (CH_3), 11.2 (CH_3); *m/z* (GC-MS) 341 (M+); Found: C, 88.08; H, 8.03; N, 4.02. Calc. for $C_{25}H_{27}N$: C, 87.93; H, 7.97; N, 4.10%.

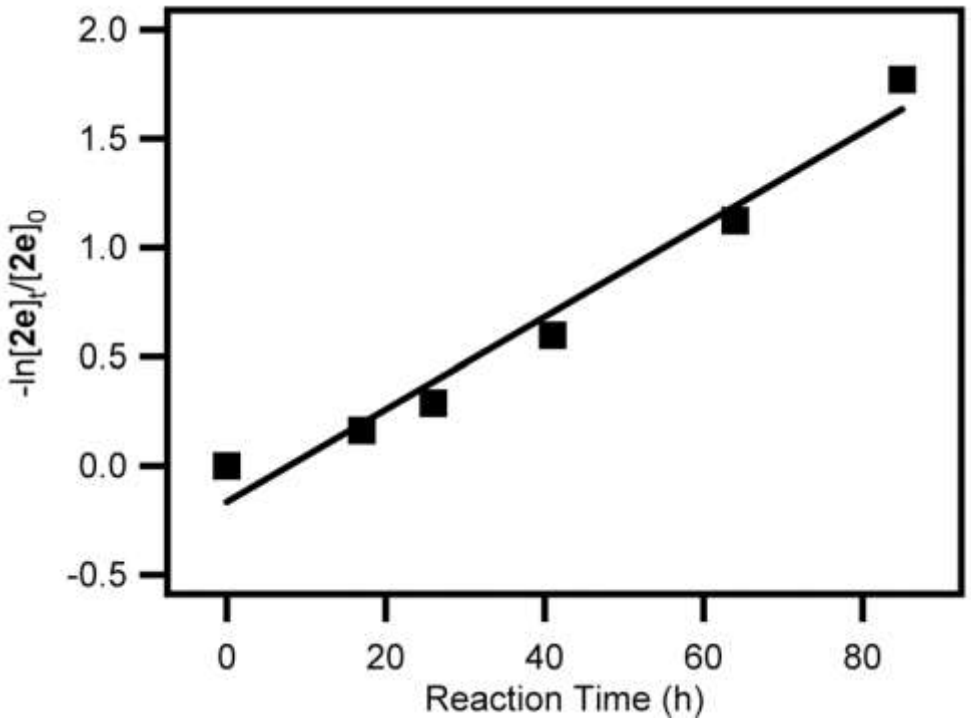
For **1e**: δ_H(400 MHz; C_6D_6) 7.68-7.01 (10 H, m, Ar), 6.24 (1 H, br s, NH), 6.19 (1H, s, C=C*H*), 2.10 (3 H, s, CH_3), 1.92 (3 H, s, CH_3), 1.89 (3 H, s, CH_3); δ_C(100 MHz; C_6D_6) 145.8, 139.3, 137.5, 134.9, 128.5, 128.4, 127.4, 127.0, 126.1, 121.2, 117.3, 114.3, 51.0 (CCH_3), 24.3 (CH_3), 12.9 (CH_3), 11.7 (CH_3); *m/z* (GC-MS) 299 (M+).

For **2e**: δ_H(400 MHz; C_6D_6) 7.68-7.01 (10 H, m, Ar), 6.44 (1 H, br s, NH), 5.53 (2H, d, *J* = 1.6 Hz, C=CH*H*), 5.27 (2 H, d, *J* = 1.6 Hz,

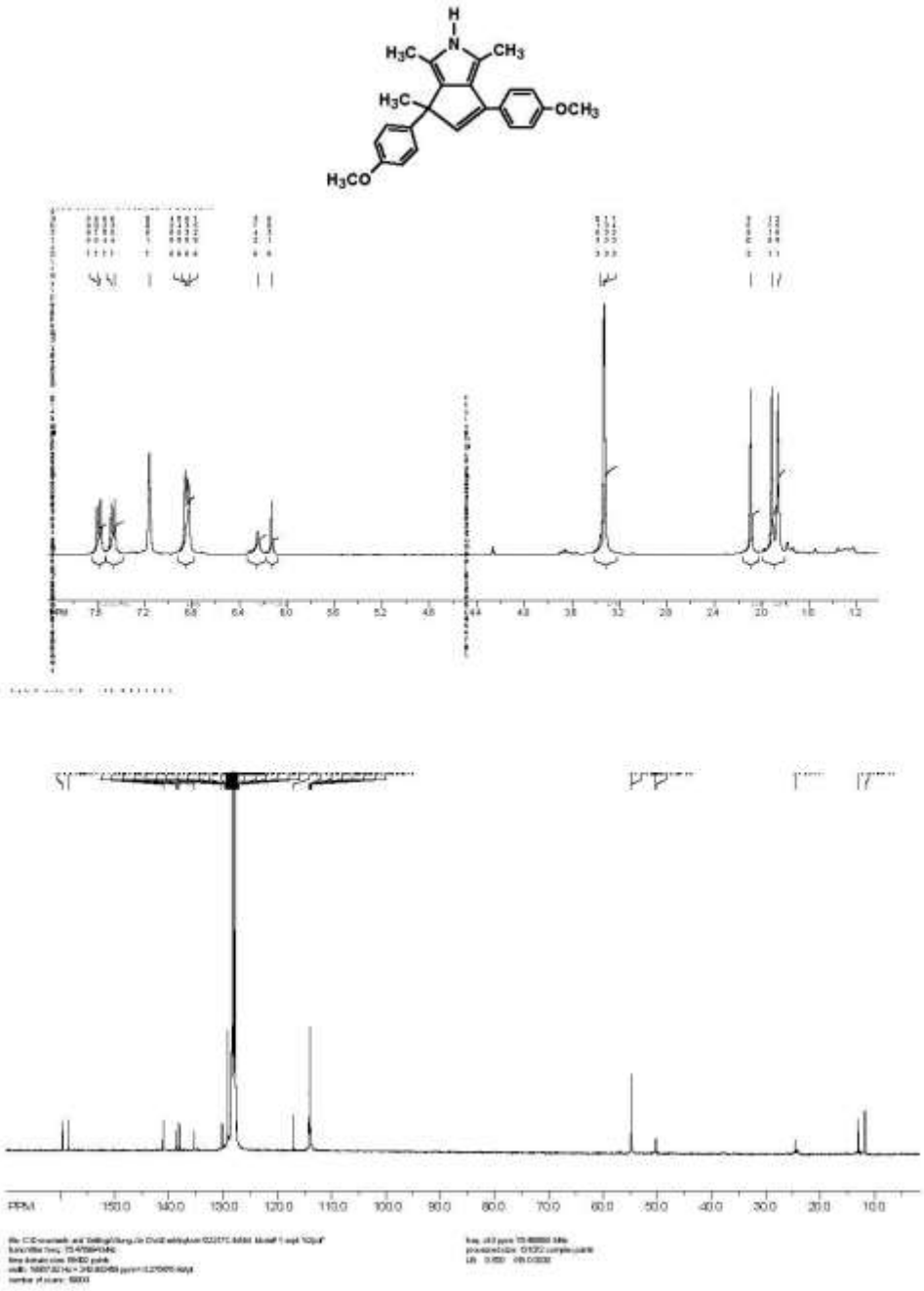
C=CH*H*), 1.92 (6 H, s, CH_3); δ_C(100 MHz, C_6D_6) 144.4, 142.9, 141.8, 128.0, 127.6, 127.2, 126.8, 123.6, 114.4, 12.0 (CH_3); *m/z* (GC-MS) 299 (M+).

For **2f**: δ_H(300 MHz; C_6D_6) 7.17-6.95 (8 H, m, Ar), 6.25 (1 H, br s, NH), 5.40 (2H, d, *J* = 2.4 Hz, C=CH*H*), 5.16 (2 H, d, *J* = 2.4 Hz, C=C*H*H), 2.07 (6 H, s, CH_3), 1.69 (6 H, s, CH_3); δH(400 MHz; acetone-d_6) 9.58 (1 H, br s, NH), 7.05-6.94 (8 H, m, Ar), 5.21 (2 H, d, *J* = 2.7 Hz, C=CH*H*), 5.98 (2 H, d, *J* = 2.7 Hz, C=C*H*H), 1.99 (6 H, s, CH_3), 1.92 (6 H, s, CH_3); δ_C(75 MHz, C_6D_6) 144.5, 143.2, 135.8, 130.6, 130.4, 126.9, 125.4, 123.4, 121.3, 117.0 (C=*C*H2), 20.6 (CH_3), 11.9 (CH_3); δ_C(100 MHz; acetone-d_6) 145.6, 143.9, 136.4, 130.0, 130.9, 127.4, 125.9, 124.6, 121.3, 116.6 (C=*C*H_2), 20.7 (CH_3), 12.0 (CH_3); *m/z* (GC-MS) 327 (M+); Found C, 87.26; H, 7.70; N, 4.32. Calc. for $C_{24}H_{25}N$: C, 88.03; H, 7.70; N, 4.28%.

For **3f**: δ_H(300 MHz; C_6D_6) 7.41-7.11 (4 H, m, Ar), 6.25 (1 H, br s, NH), 5.91 (1H, s, C=C*H*), 5.51 (1 H, d, *J* = 1.8 Hz, C=CH*H*), 5.05 (1 H, d, *J* = 1.8 Hz, C=C*H*H), 2.27 (3 H, s, CH_3), 1.81 (3 H, s, CH_3), 1.64 (3 H, s, CH_3); δ_C(75 MHz, C_6D_6) 145.9, 144.3, 136.1, 130.1, 129.8, 125.9, 124.6, 123.3, 121.0, 20.1 (CH_3), 12.7 (CH_3), 12.6 (CH_3); *m/z* (GC-MS) 211 (M+).

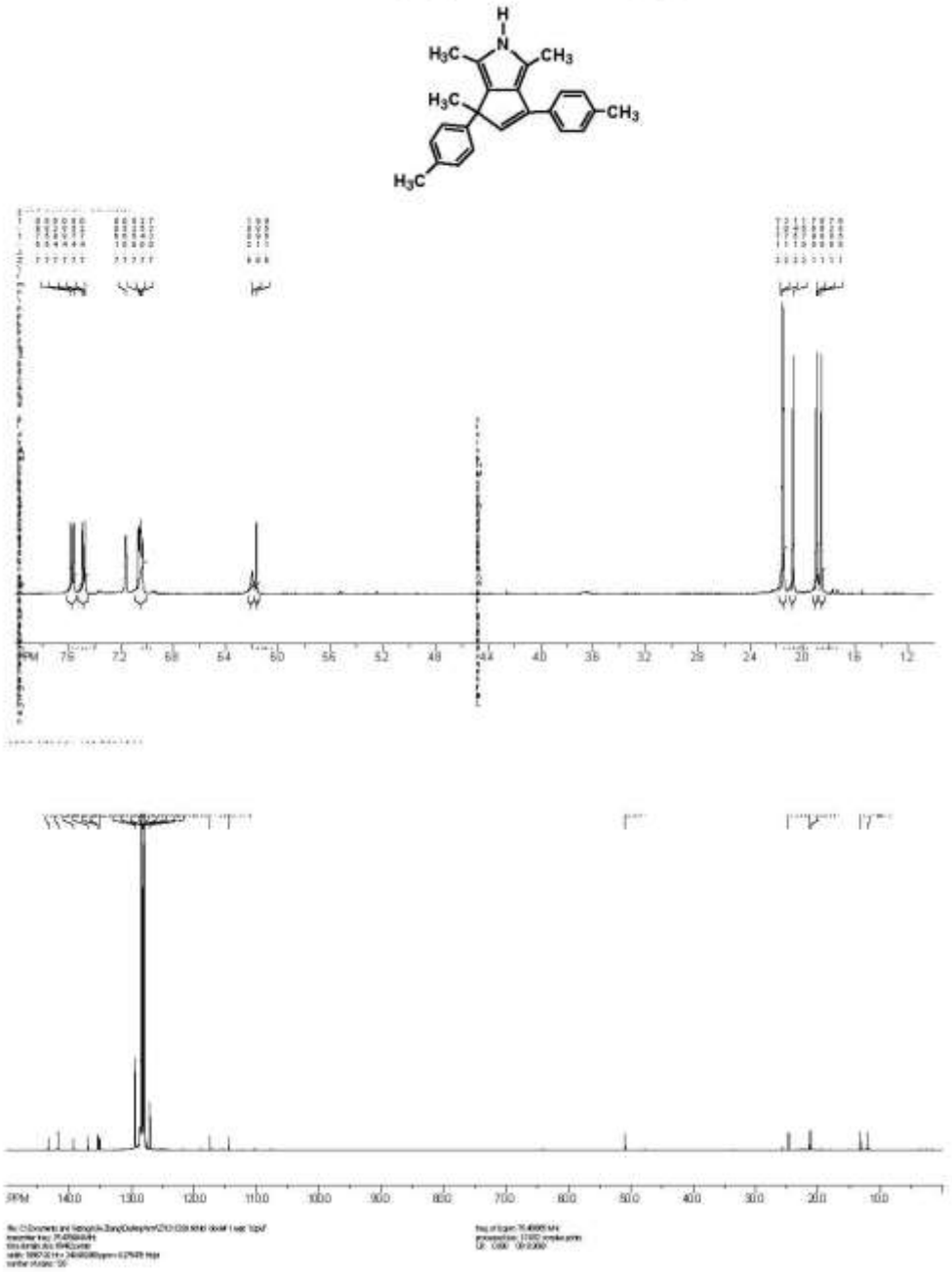


^{1}H and $^{13}C\{^{1}H\}$ NMR of **1a** in C_6D_6



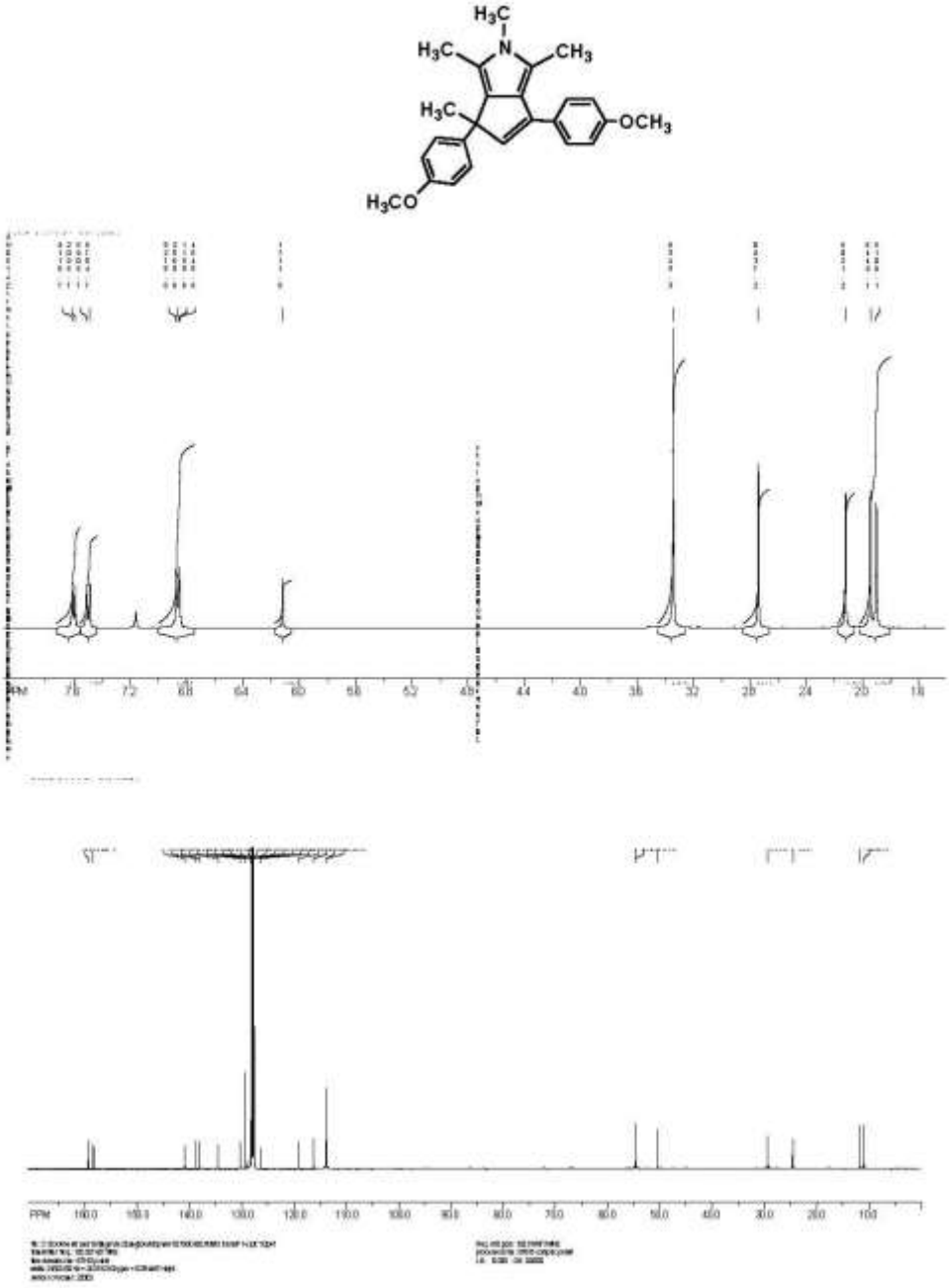
Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.

^{1}H and $^{13}C\{^{1}H\}$ NMR of **1b** in C_6D_6



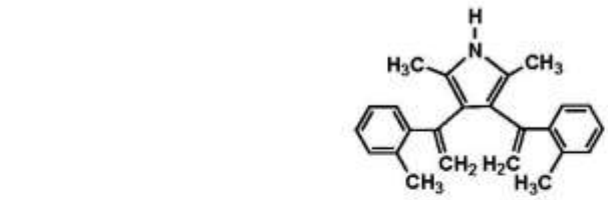
Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.

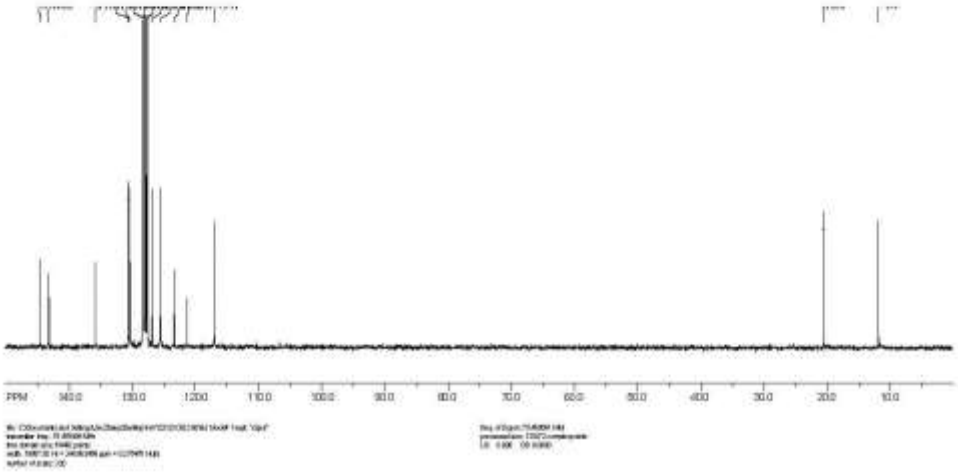
^{1}H and $^{13}C\{^{1}H\}$ NMR of **1c** in C_6D_6



Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.

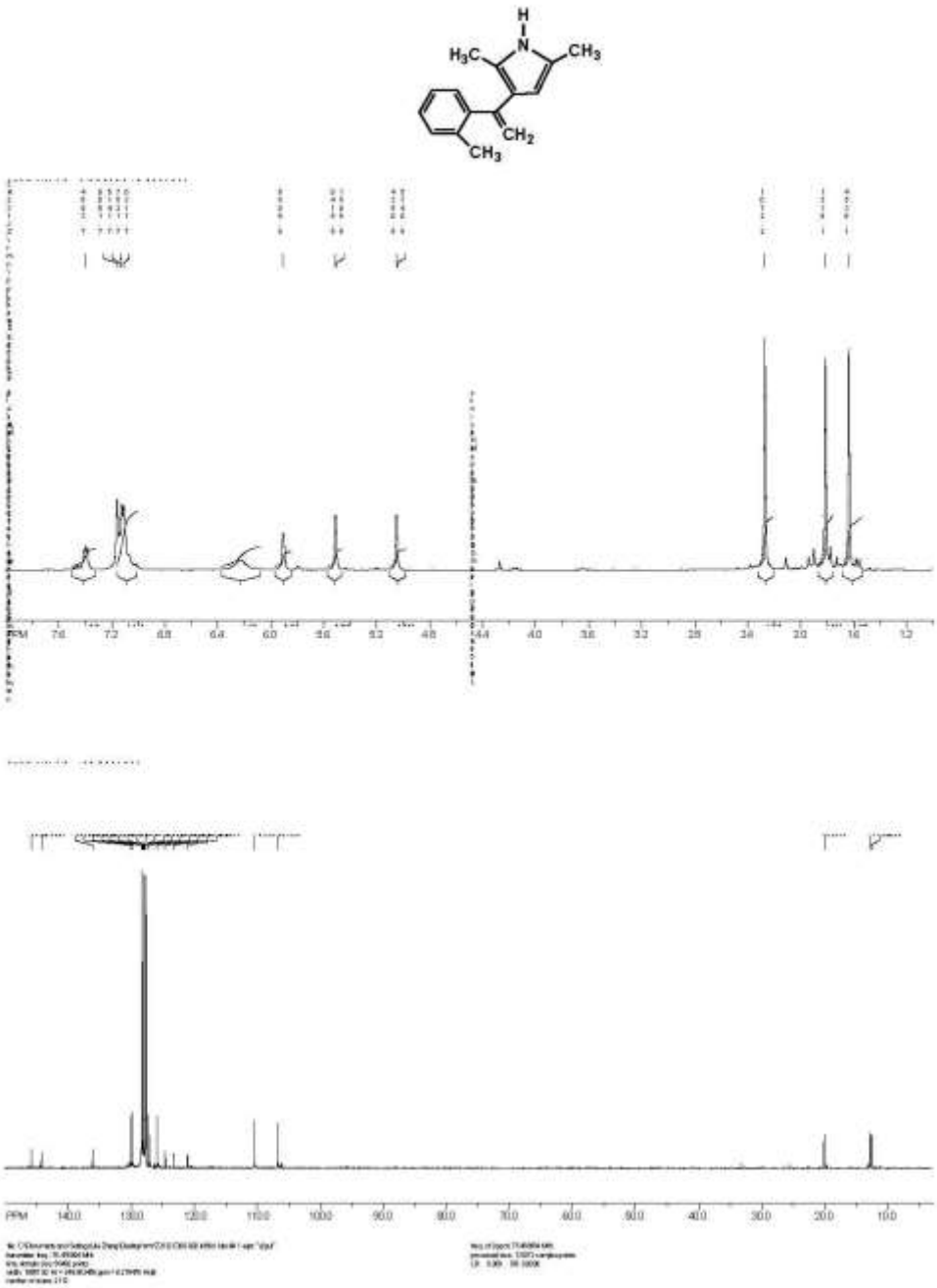
$^{13}C\{^{1}H\}$ NMR of **2f** in C_6D_6





Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.

^{1}H and $^{13}C\{^{1}H\}$ NMR of **3f** in C_6D_6



Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.

Acknowledgments: Financial support from the US National Institute of Health General Medical Sciences (R15 GM55987) is gratefully acknowledged.

Footnotes: †Electronic supplementary information (ESI) available: The experimental procedure and characterization data of organic products. See DOI: 10.1039/b804263b

‡*Representative experimental procedure:* In a glove box, $Ru_3(CO)_{12}$ (0.03 mmol), NH_4PF_6 (0.1 mmol), 2,5-dimethylpyrrole (1.0 mmol) and an alkyne (2.0 mmol) were dissolved in benzene (5 mL) in a medium-walled 25 mL Schlenk tube, equipped with a Teflon stopcock and a magnetic stirring bar. The reaction tube was sealed, brought out of the box and heated in an oil bath at 95 °C for 36–48 h. The tube was opened to air at room temperature and the crude product mixture analysed by GC. The solvent was removed using a rotary evaporator and the organic product was isolated by column chromatography on silica gel (hexane/CH_2Cl_2) under a nitrogen atmosphere.

For **1b**: δ_H(400 MHz; C_6D_6) 7.58–7.03 (8 H, m, Ar), 6.20 (1 H, br s, NH), 6.17 (1 H, s, CQCH), 2.15 (6 H, s, CH_3), 2.08 (3 H, s, CH_3), 1.90 (3 H, s, CH_3) and 1.86 (3 H, s, CH_3); δ_C(75 MHz; C_6D_6) 142.9, 141.4, 138.9, 136.6, 135.1, 135.0, 134.7, 129.1, 127.9, 126.7, 117.0, 114.1, 50.6 (*C*CH_3), 24.3 (CH_3), 21.1 (CH_3), 20.9 (CH_3), 12.8 (CH_3) and 11.6 (CH_3); *m/z* (GC-MS) 327 (M^+); Found: C, 88.02; H, 7.62; N, 4.31.

Notes and references

1. For recent reviews, see: (a) C. Jia, T. Kitamura and Y. Fujiwara, Acc. Chem. Res., 2001, 34, 633; (b) F. Kakiuchi and S. Murai, Acc. Chem. Res., 2002, 35, 826; (c) D. Alberico, M. E. Scott and M. Lautens, Chem. Rev., 2007, 107, 174.
2. Recent selected examples: (a) W. Lu, C. Jia, T. Kitamura and Y. Fujiwara, Org. Lett., 2000, 2, 2927; (b) R. K. Thalji, K. A. Ahrendt, R. G. Bergman and J. A. Ellman, J. Am. Chem. Soc., 2001, 123, 9692; (c) C. Liu, X. Han, X. Wang and R. A. Widenhoefer, J. Am. Chem. Soc., 2004, 126, 3700; (d) X.Wan, Z. Ma, B. Li, K. Zhang, S. Cao, S. Zhang and Z. Shi, J. Am. Chem. Soc., 2006, 128, 7416; (e) X. Chen, C. E. Goodhue and J.-Q. Yu, J. Am. Chem. Soc., 2006, 128, 12634; (f) J. C. Lewis, R. G. Bergman and J. A. Ellman, J. Am. Chem. Soc., 2007, 129, 5332.
3 (a) D. Kalyani, N. R. Deprez, L. V. Desai and M. S. Sanford, J. Am. Chem. Soc., 2005, 127, 7330; (b) H.-Y. Thu, W.-Y. Yu and C.-M. Che, J. Am. Chem. Soc., 2006, 128, 9048; (c) B.-J. Li, S.-L. Tian, Z. Fang and Z.-J. Shi, Angew. Chem., Int. Ed., 2008, 47, 1115.
4 E. M. Ferreira and B. M. Stoltz, J. Am. Chem. Soc., 2003, 125, 9578.

5 (a) C. S. Yi, S. Y. Yun and I. A. Guzei, J. Am. Chem. Soc., 2005, 127, 5782; (b) C. S. Yi and S. Y. Yun, J. Am. Chem. Soc., 2005, 127, 17000.
6 (a) C. Jones, D. Taube, V. R. Ziatdinov, R. A. Periana, R. J. Nielsen, J. Oxgaard and W. A. Goddard III, Angew. Chem., Int. Ed., 2004, 43, 4626; (b) Z. Li, D. A. Capretto, R. O. Rahaman and C. He, J. Am. Chem. Soc., 2007, 129, 12058.

About the Authors: Department of Chemistry, Marquette University, Milwaukee, Wisconsin, USA. Fax: +1 414-288-7066; Tel: +1 414-288-3536 Corresponding author. Chae S. Yi: chae.yi@marquette.edu

Chemical Communications, Vol. 20, (May 2008): pg. 2349-2351. DOI.